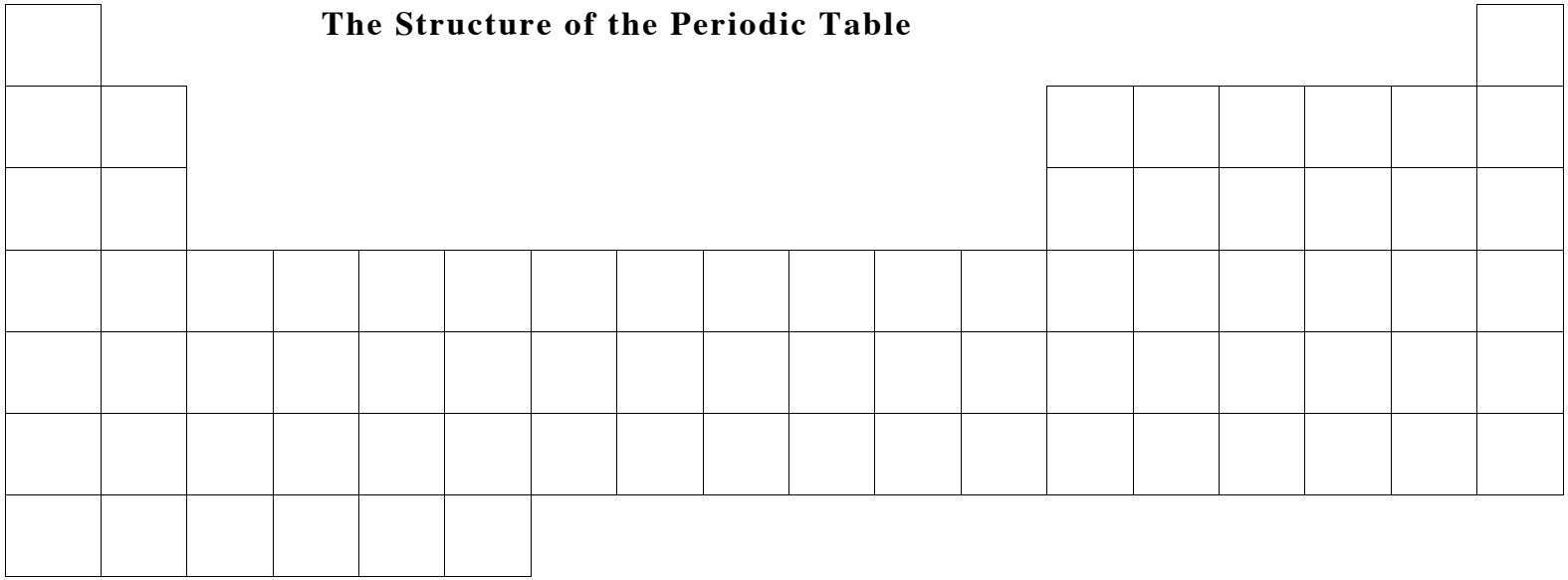
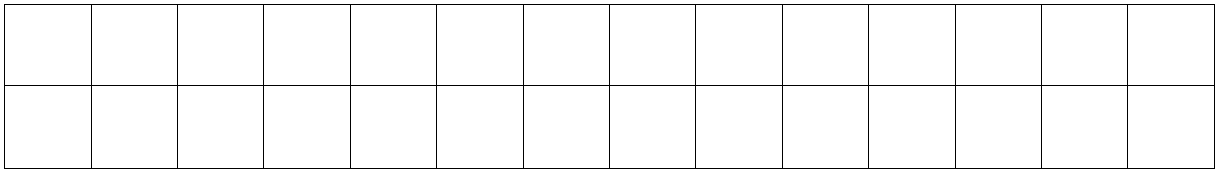
The Structure of the Periodic Table





Directions:

1. Label the columns with the both the 1 – 18 and IA – VIIIA designations.
2. Write the period numbers in the left margin.
3. Draw a heavy line between the metals and the nonmetals.
4. Color the groups listed in the key with different colored pencils, indicating colors in the key.
5. Write the atomic numbers in the upper left hand corner of elements 57, 58, 71, 72, 89, 90, 103, and 104. Write the symbols of lanthanum and actinium in the correct boxes.
6. Write the symbols of the first twenty elements and {Ag, As, Au, Ba, Br, Co, Cr, Cu, Cs, Fe, Fr, Hg, I, Mn, Ni, Pb, Pt, Pu, Se, Sn, Sr, Ti, U, Zn} in the proper locations.
7. Locate each of the 7 elements that exist as diatomic molecules. Label in a way that you will remember them.
8. Mark the locations of the metalloids with an "X" in the corner of their squares.
9. Circle the elements C, H, O, N (common in organic and biochem)

K E Y

- ☐ alkali metals
- ☐ alkaline earth metals
- ☐ halogen family
- ☐ noble gases
- ☐ transition metals
- ☐ lanthanide series
- ☐ actinide series